|Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_| |Class: XI | | Paper: Chemistry | | Time: 60 minutes | |Max. Marks: 15 | |Test # 1|

|NAME:| \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | |F.NAME:| \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_|

KQS COACHING CENTER

Q1: Solve the following

i. 15.5m, 651.8cm, and 4291m ii.

Q2: Define (i) Molecular Mass (ii) Molecular Formula (iii) Avogadro’s number (iv) Mole.

Q3: Calculate the mass of Carbondioxide (CO­­2­) that can be obtained by heating 30g of limestone (CaCO3) and also calculate the mass of Calcium Oxide (CaO) . CaCO3  - - - - - - - - - - > CaO + CO2

Q4: Calculate the volume of CO2 gas produced at Standard Temperature and pressure by the combustion of 40g of CH4

CH4  + 2O2 - - - - - - - > CO2  + 2H2O

Q5: What volume of O2  at S.T.P is required to burn 500dm3 of Ethene C2H2 gas ? What volume of CO2 will be formed?

C2H4 + 3O2 - - - - - - - - > 2 CO2 + 2H2O

Q6: i. Calculate the number of atoms in 9.2g of Na (Na = 23 a.m.u.).

ii. Calculate the mass in grams of 3.01x1020molecules of glucose (C6 H12 O6).

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